Name _____ Date _____

Gram to Particles

Calculate the following conversions:
Show your work:

1. 12.8 g C_9H_{20} = _____ C_9H_{20} molecules

2. 129 g C_6H_{14} = _____ C_6H_{14} molecules

3. 360 g C_5H_{12} = _____ C_5H_{12} molecules

4. 470 g Cu_2CO_3 = _____ Cu_2CO_3 formula units

5. 132 g C = _____ C atoms

6. 87 g NaCl = _____ NaCl formula units

7. 100g C_7H_{16} = _____ C_7H_{16} molecules

8. 217g Na_2O = _____ Na_2O formula units

9. 43 g Pb^{2+} = _____ Pb^{2+} ions

10. 100 g Na_2CO_3 = _____ Na_2CO_3 formula units

11. 113.6 g $C_{10}H_{22}$ = _____ $C_{10}H_{22}$ molecules

Particles to Grams

12. 6.02×10^{23} formula units $MgCl_2$ = _____ g $MgCl_2$

13. 9.03×10^{23} molecules CH_4 = _____ g CH_4

14. 1.806×10^{24} molecules H_2O = _____ g H_2O

15. 3.612×10^{24} atoms Na = _____ g Na

16. 1.204×10^{24} molecules O_2 = _____ g O_2

17. 7.224×10^{24} molecules C_4H_{10} = _____ g C_4H_{10}

18. 6.02×10^{23} formula units $Pb(NO_3)_3$ = _____ gl $Pb(NO_3)_3$

19. 1.505×10^{24} Ca^{+2} ions = _____ g Ca^{+2}

20. 2.107×10^{24} formula units KCl = _____ g KCl

21. 2.709×10^{24} Cl^{-1} ions = _____ g Cl^{-1}

22. 3.01×10^{23} molecules Cl_2 = _____ g Cl_2