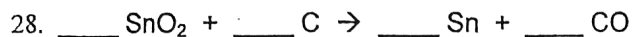


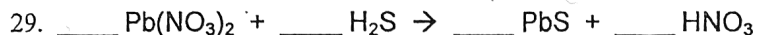
Chemistry: Balancing Chemical Equations

Directions: First, balance each of the chemical equations below. Then, classify each reaction as **synthesis, decomposition, single-replacement, double-replacement, or combustion.**

Balance the equation...

...and classify it.

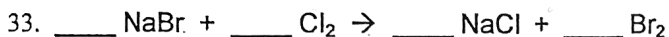


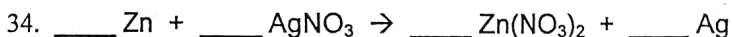






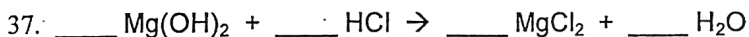




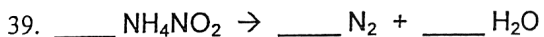


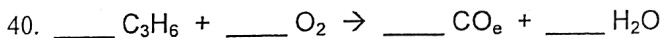




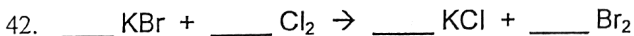


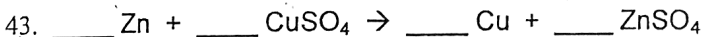


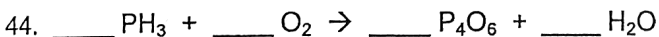


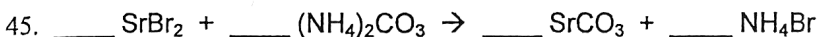


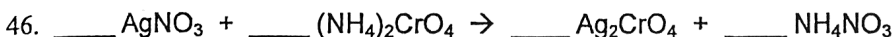


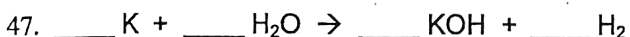


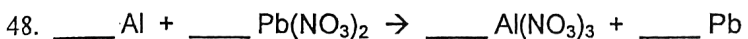










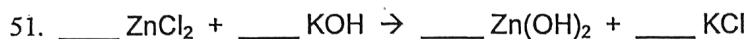


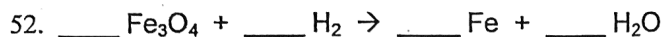
Balance the equation...

Name: _____
...and classify it.















Chemical Word Equations

Directions: Write a balanced chemical equation for each of the word equations below.

56. Solid aluminum hydroxide reacts with aqueous nitric acid (HNO₃) to yield soluble aluminum nitrate and liquid water

57. Aqueous lead (IV) nitrate reacts with aqueous sodium sulfate to yield a lead (IV) sulfate precipitate and soluble sodium nitrate

58. Aqueous sodium hydroxide reacts with carbon dioxide gas to yield soluble sodium carbonate and liquid water

59. Solid magnesium oxide reacts with hydrochloric acid solution (HCl) to yield a solution of magnesium chloride and liquid water

60. Solid zinc metal reacts with dissolved sulfuric acid (H₂SO₄) to yield aqueous zinc sulfate and hydrogen gas

61. Solid iron (III) oxide reacts with solid aluminum metal to yield solid aluminum oxide and solid iron metal.

62. When dissolved beryllium chloride reacts with dissolved silver nitrate in water, aqueous beryllium nitrate and silver chloride powder are made.

63. When liquid isopropanol (rubbing alcohol = C₃H₈O) burns in oxygen, carbon dioxide, water, and heat are produced.